

Ácido	Fórmula y ecuación de ionización		$K_a$	$pK_a$
Hipocloroso	HOCl	$\rightleftharpoons H^+ + OCl^-$	$3.5 \times 10^{-8}$	7.46
Hipoyodoso	HOI	$\rightleftharpoons H^+ + OI^-$	$2.3 \times 10^{-11}$	10.64
Nitroso	HNO <sub>2</sub>	$\rightleftharpoons H^+ + NO_2^-$	$4.5 \times 10^{-4}$	3.35
Oxálico	(COOH) <sub>2</sub>	$\rightleftharpoons H^+ + COOCOOH^-$	$5.9 \times 10^{-2} = K_{a1}$	1.23
	COOCOOH <sup>-</sup>	$\rightleftharpoons H^+ + (COO)_2^{2-}$	$6.4 \times 10^{-5} = K_{a2}$	4.19
Fenol	HC <sub>6</sub> H <sub>5</sub> O	$\rightleftharpoons H^+ + C_6H_5O^-$	$1.3 \times 10^{-10}$	9.89
Fosfórico	H <sub>3</sub> PO <sub>4</sub>	$\rightleftharpoons H^+ + H_2PO_4^-$	$7.5 \times 10^{-3} = K_{a1}$	2.12
	H <sub>2</sub> PO <sub>4</sub> <sup>-</sup>	$\rightleftharpoons H^+ + HPO_4^{2-}$	$6.2 \times 10^{-8} = K_{a2}$	7.21
	HPO <sub>4</sub> <sup>2-</sup>	$\rightleftharpoons H^+ + PO_4^{3-}$	$3.6 \times 10^{-13} = K_{a3}$	12.44
Fosforoso	H <sub>3</sub> PO <sub>3</sub>	$\rightleftharpoons H^+ + H_2PO_3^-$	$1.6 \times 10^{-2} = K_{a1}$	1.80
	H <sub>2</sub> PO <sub>3</sub> <sup>-</sup>	$\rightleftharpoons H^+ + HPO_3^{2-}$	$7.0 \times 10^{-7} = K_{a2}$	6.15
Selénico	H <sub>2</sub> SeO <sub>4</sub>	$\rightleftharpoons H^+ + HSeO_4^-$	Muy grande = $K_{a1}$	
	HSeO <sub>4</sub> <sup>-</sup>	$\rightleftharpoons H^+ + SeO_4^{2-}$	$1.2 \times 10^{-2} = K_{a2}$	1.92
Selenioso	H <sub>2</sub> SeO <sub>3</sub>	$\rightleftharpoons H^+ + HSeO_3^-$	$2.7 \times 10^{-3} = K_{a1}$	2.57
	HSeO <sub>3</sub> <sup>-</sup>	$\rightleftharpoons H^+ + SeO_3^{2-}$	$2.5 \times 10^{-7} = K_{a2}$	6.60
Sulfúrico	H <sub>2</sub> SO <sub>4</sub>	$\rightleftharpoons H^+ + HSO_4^-$	Muy grande = $K_{a1}$	
	HSO <sub>4</sub> <sup>-</sup>	$\rightleftharpoons H^+ + SO_4^{2-}$	$1.2 \times 10^{-2} = K_{a2}$	1.92
Sulfuroso	H <sub>2</sub> SO <sub>3</sub>	$\rightleftharpoons H^+ + HSO_3^-$	$1.2 \times 10^{-2} = K_{a1}$	1.92
	HSO <sub>3</sub> <sup>-</sup>	$\rightleftharpoons H^+ + SO_3^{2-}$	$6.2 \times 10^{-8} = K_{a2}$	7.21
Teluroso	H <sub>2</sub> TeO <sub>3</sub>	$\rightleftharpoons H^+ + HTeO_3^-$	$2 \times 10^{-3} = K_{a1}$	2.70
	HTeO <sub>3</sub> <sup>-</sup>	$\rightleftharpoons H^+ + TeO_3^{2-}$	$1 \times 10^{-8} = K_{a2}$	8.00