

| Ácido | Fórmula y ecuación de ionización | | K_a | pK_a |
|-----------------------|---|---|--------------------------------|--------|
| Acético | CH_3COOH | $\rightleftharpoons \text{H}^+ + \text{CH}_3\text{COO}^-$ | 1.8×10^{-5} | 4.74 |
| Arsénico | H_3AsO_4 | $\rightleftharpoons \text{H}^+ + \text{H}_2\text{AsO}_4^-$ | $2.5 \times 10^{-4} = K_{a1}$ | 3.60 |
| | H_2AsO_4^- | $\rightleftharpoons \text{H}^+ + \text{HASO}_4^{2-}$ | $5.6 \times 10^{-8} = K_{a2}$ | 7.25 |
| | HASO_4^{2-} | $\rightleftharpoons \text{H}^+ + \text{AsO}_4^{3-}$ | $3.0 \times 10^{-13} = K_{a3}$ | 12.52 |
| Arsenioso | H_3AsO_3 | $\rightleftharpoons \text{H}^+ + \text{H}_2\text{AsO}_3^-$ | $6.0 \times 10^{-10} = K_{a1}$ | 9.22 |
| | H_2AsO_3^- | $\rightleftharpoons \text{H}^+ + \text{HASO}_3^{2-}$ | $3.0 \times 10^{-14} = K_{a2}$ | 13.52 |
| Benzoico | $\text{C}_6\text{H}_5\text{COOH}$ | $\rightleftharpoons \text{H}^+ + \text{C}_6\text{H}_5\text{COO}^-$ | 6.3×10^{-5} | 4.20 |
| Bórico* | B(OH)_3 | $\rightleftharpoons \text{H}^+ + \text{BO(OH)}_2^-$ | $7.3 \times 10^{-10} = K_{a1}$ | 9.14 |
| | BO(OH)_2^- | $\rightleftharpoons \text{H}^+ + \text{BO}_2(\text{OH})^{2-}$ | $1.8 \times 10^{-13} = K_{a2}$ | 12.74 |
| | $\text{BO}_2(\text{OH})^{2-}$ | $\rightleftharpoons \text{H}^+ + \text{BO}_3^{3-}$ | $1.6 \times 10^{-14} = K_{a3}$ | 13.80 |
| Carbónico | H_2CO_3 | $\rightleftharpoons \text{H}^+ + \text{HCO}_3^-$ | $4.2 \times 10^{-7} = K_{a1}$ | 6.38 |
| | HCO_3^- | $\rightleftharpoons \text{H}^+ + \text{CO}_3^{2-}$ | $4.8 \times 10^{-11} = K_{a2}$ | 10.32 |
| Cítrico | $\text{C}_3\text{H}_5\text{O}(\text{COOH})_3$ | $\rightleftharpoons \text{H}^+ + \text{C}_4\text{H}_5\text{O}_3(\text{COOH})_2^-$ | $7.4 \times 10^{-3} = K_{a1}$ | 2.13 |
| | $\text{C}_4\text{H}_5\text{O}_3(\text{COOH})_2^-$ | $\rightleftharpoons \text{H}^+ + \text{C}_5\text{H}_5\text{O}_5\text{COOH}^{2-}$ | $1.7 \times 10^{-5} = K_{a2}$ | 4.77 |
| | $\text{C}_5\text{H}_5\text{O}_5\text{COOH}^{2-}$ | $\rightleftharpoons \text{H}^+ + \text{C}_6\text{H}_5\text{O}_7^{3-}$ | $7.4 \times 10^{-7} = K_{a3}$ | 6.13 |
| Ciánico | HOCN | $\rightleftharpoons \text{H}^+ + \text{OCN}^-$ | 3.5×10^{-4} | 3.46 |
| Fórmico | HCOOH | $\rightleftharpoons \text{H}^+ + \text{HCOO}^-$ | 1.8×10^{-4} | 3.74 |
| Hidrazoico | HN_3 | $\rightleftharpoons \text{H}^+ + \text{N}_3^-$ | 1.9×10^{-5} | 4.72 |
| Hidrocianico | HCN | $\rightleftharpoons \text{H}^+ + \text{CN}^-$ | 4.0×10^{-10} | 9.40 |
| Hidroflórico | HF | $\rightleftharpoons \text{H}^+ + \text{F}^-$ | 7.2×10^{-4} | 3.14 |
| Peróxido de hidrógeno | H_2O_2 | $\rightleftharpoons \text{H}^+ + \text{HO}_2^-$ | 2.4×10^{-12} | 11.62 |
| Hidrosulfúrico | H_2S | $\rightleftharpoons \text{H}^+ + \text{HS}^-$ | $1.0 \times 10^{-7} = K_{a1}$ | 7.00 |
| | HS^- | $\rightleftharpoons \text{H}^+ + \text{S}^{2-}$ | $1.0 \times 10^{-19} = K_{a2}$ | 19.00 |
| Hipobromoso | HOBr | $\rightleftharpoons \text{H}^+ + \text{OBr}^-$ | 2.5×10^{-9} | 8.60 |